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**Photodegradation of Polyimides. VI. Effect of Donor-Acceptor Groups on the
Photooxidative Stability of Polyimides and Model Compounds**

by

C.E. Hoyle, D. Creed, P. Subramanian, R. Nagarajan, C. Pandey and E.T. Anzures

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**Photodegradation of Polyimides. Part VI. Effect
of Donor-Acceptor Groups on the Photooxidative
Stability of Polyimides and Model Compounds**

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ABSTRACT: The photophysics of polyimides and several N-arylphthalimide model compounds has been investigated as part of an effort to understand the factors that control the oxidative photodegradation of these materials. Substituents that increase the donor-acceptor (D-A) character of both polymers and models cause an enhancement of the red-shifted electronic absorption and a shift to lower energy of the broad, weak fluorescence of these materials. Increasing solvent polarity causes a similar effect on the absorption and fluorescence spectra of the model compounds. Fluorescence quantum yields decrease with increasing D-A character in the series of model compounds, and with increasing solvent polarity for a given compound. The triplet states of both polymers and models have been characterized in solution by laser flash photolysis, and intersystem crossing quantum yields have been determined. The triplet yields also decrease with increasing D-A character of the chromophore and with increasing solvent polarity. A comparison of fluorescence and phosphorescence spectra reveals a very small singlet-triplet energy gap. It seems likely that the lowering, by increased D-A character, of the yields of the triplet states that are intermediates in the photooxidative degradation, enhances the photostability of aromatic based polyimides.

Introduction

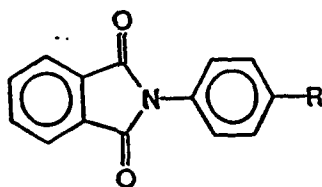
Aromatic polyimides are important high performance materials because of their superior thermal stability, excellent mechanical properties, and low dielectric constants. There are many reports of the thermal behavior of high temperature, heat resistant polyimides but relatively little is known about their photochemical behavior. We have recently begun a comprehensive study¹⁻⁵ of the photochemistry and photophysics of a series of polyimides exemplified by 6F-ODA and 6F-6F (see structures and appropriate designations), derived by reaction of a bis-aromatic dianhydride, 6F, in which two phthalic anhydride moieties are linked by a hexafluoroisopropylidene 'hinge' group, with bridged (either oxygen or hexafluoroisopropylidene) aromatic diamines, and by PMDA-ODA, derived from pyromellitic dianhydride and 4,4'-oxydianiline. In the course of this investigation, we also conducted a comprehensive study of a series of N-arylphthalimides (see structures designated PA-A, PA-POA, PA-CIA, PA-MOA and PA-CNA) with differing substituents at the *para*-position of the N-aryl ring^{4,5}. These phthalimides model (*vide infra*) many of the photophysical and photochemical properties of the much less tractable polymers.

Broad band mercury lamp irradiation of polymers such as 6F-ODA, 6F-6F, and 6F-6H in the presence of oxygen results in their complete degradation to small molecule fragments. For example, 6F-ODA, when

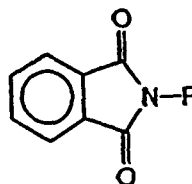
irradiated in the presence of oxygen, yields⁴ photoproducts such as benzoic acid, H₂O, CO, CO₂, and CF₃H : Irradiation of films of 6F-ODA and 6F-6F results in progressive weight loss by clean ablation of the surface. In contrast, irradiation of 6F-ODA through a Pyrex filter results in chain breaks but no facile breakdown to small molecule fragments or surface ablation. Upon examination of model systems, we found that irradiation of model N-arylphthalimides through Pyrex in the presence of oxygen affords phthalic anhydride as the major photoproduct.⁴ Phthalic anhydride undergoes secondary photolysis to unknown photoproducts only under broad-band mercury lamp irradiation, which includes light of wavelength less than 300 nm. These observations have led us to propose⁴ a mechanism for the extensive photooxidative degradation of fluorinated polyimides involving initial production of aryl anhydride functional groups in the irradiated polymer followed by their subsequent photolysis. We have obtained evidence for the first part of this mechanism by direct observation⁵ of aryl anhydride functional groups (by their characteristic IR bands) in 6F-6F films irradiated in the presence of air.

It is clear from our previous work on photooxidative degradation of both polymers and models that the least photolabile materials are those in which the donor-acceptor (D-A) character of the chromophore is predicted to be the most pronounced. The order of stability of the polymers (as determined by weight loss) is PMDA-ODA > 6F-ODA > 6F-6F, and, for the

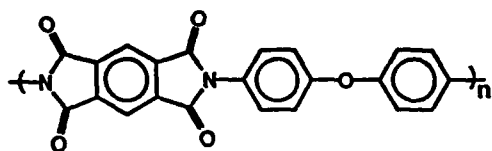
models (as determined by degradation quantum yields) is PA-POA > PA-A \approx PA-CIA > PA-CNA. Presumably, the pyromellitimide moiety in PMDA-ODA is a more powerful electron acceptor than the 6F moiety in 6F-ODA, and the 6F bridging group in the diamine moiety of 6F-6F suppresses the D-A character of 6F-6F relative to 6F-ODA. The photostability trend for the N-arylphthalimide models parallels the expected electron donating ability of the *para*-substituent, i.e., OPh > H > CN. We have also obtained evidence from quenching experiments that the triplet state is a key intermediate in PA-A photodegradation. In this paper we will report details of D-A effect on the singlet states of 6F based polymers and appropriate model phthalimides, direct observation of the triplet states of these materials in fluid solution, measurements of intersystem crossing quantum yields (Φ_{isc}) and, finally, low temperature emission experiments that enable estimation of singlet-triplet energy differences. Our results suggest that D-A stabilization of the lowest singlet state of both polymers and models reduces photochemical reactivity, to a certain extent, by reduction of intersystem crossing yields.



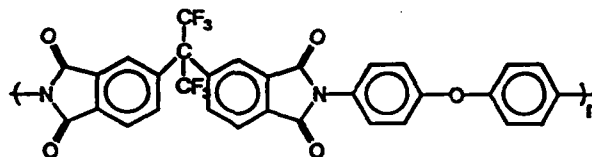
R = CN, PA-CNA
 R = H, PA-A
 R = Cl, PA-CIA
 R = OPh, PA-POA
 R = OMe, PA-MOA



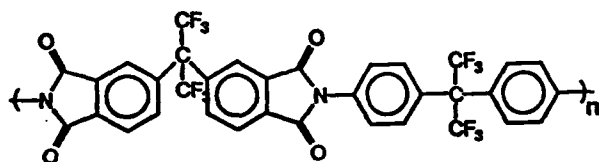
R = Me, PA-Me
 R = Pr, PA-Pr



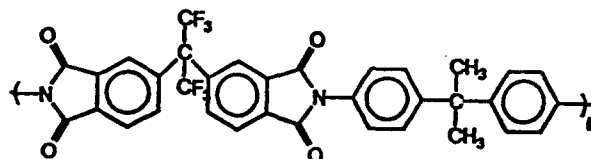
PMDA-ODA



6F-ODA



6F-6F



6F-6H

Experimental Section

The synthesis and purification of both polymers and model compounds has been described elsewhere.^{1,4} All solvents were obtained from Burdick and Jackson and used as received. UV-VIS absorption spectra were recorded on a Perkin-Elmer Lambda 6 spectrophotometer and fluorescence spectra on a Spex Fluorolog-2 spectrofluorometer. A Spex phosphorescence attachment was used to obtain phosphorescence spectra at different delay times after pulsed excitation. Fluorescence quantum yields were obtained relative to anthracene ($\Phi_F = 0.3$) in cyclohexane⁶. Laser flash photolysis experiments were carried out using a Lumonics HyperEx 440 excimer laser as an excitation source operating at 248 nm (KrF) or 351 nm

(XeF), and an Applied Photophysics (AP) pulsed Xenon lamp as monitoring source, the output from which was analyzed using an AP monochromator/fast PMT system as reported earlier⁵. Extinction coefficients of the triplet states were determined by the energy transfer method⁷ from benzophenone. Intersystem crossing yields, Φ_{ISC} , were determined relative to benzophenone ($\Phi_{ISC} = 1.0$) by laser flash photolysis at 248 nm.

Results and Discussion

A. Fluorinated Polyimides : Photolysis Dependence on Amine Donor Characteristics

The absorption and fluorescence spectra of the polyimides designated 6F-6F, 6F-ODA and 6F-6H are similar to those previously reported⁸ for related materials. UV-VIS absorption spectra of thin films of these polymers cured at 300°C are shown in Fig. 1 with the 6F-ODA film being considerably red shifted compared to 6F-6F and 6F-6H. The red shift in the absorption spectra occurs in accordance with expected increasing D-A interactions, i.e., 6F-6F < 6F-6H < 6F-ODA. Solutions and thin films of the 6F based polymers exhibit a broad, weak fluorescence maximum. For example, 6F-6F has a maximum at 542 nm in acetonitrile, 530 nm in tetrahydrofuran, and 525 nm in CH₂Cl₂. The results are typical for the 6F based polymers, all of which exhibit a red shift in the fluorescence peak maximum with increasing solvent polarity. Of particular interest in the

present case is to compare the shift in the fluorescence peak maximum in a common solvent for 6F-6F, 6F-6H and 6F-ODA. The entries in Table I show fluorescence emission maxima for all three polymers in at least two solvents : The shift in the peak maximum to lower energies with the expected order of increasing D-A interaction ($6F-6F < 6F-6H < 6F-ODA$) occurs in each solvent. In addition, the intensity of the triplet spectrum of all three polymers in CH_2Cl_2 suggests that the intersystem crossing yield to the triplet state of the aryl imide is greatest for 6F-6F and least for 6F-ODA with 6F-6H being intermediate. This assumes that the extinction coefficients of the triplet states are approximately equal. Based on the donor character expected for the amine based component of the aryl imide repeat unit ($ODA > 6H > 6F$), we might expect to see the photolytic decomposition, which we will demonstrate in a subsequent section in this paper, to proceed, at least to an extent, from the triplet state, to follow the same order, i.e., 6F-6F would be the more photolabile polymer. In accordance with this prediction, photolysis of 6F-6F films led to very rapid decomposition compared to 6F-6H and 6F-ODA films. The rapid decomposition of 6F-6F is even more surprising when we consider that its UV-VIS absorption (not shown) in the range of the photolysis source (unfiltered medium pressure Hg lamp) is much less than for 6F-6H and 6F-ODA. Fig. 2 shows changes in the molecular weight at the peak maximum for 6F-ODA, 6F-6F, and 6F-6H as a function of photolysis time in air. As indicated above, the order of

decomposition is 6F-6F > 6F-6H > 6F-ODA. Obviously the photostability of the films parallels the D-A acceptor properties of the photoreactive repeat unit and the apparent intersystem crossing yield to the triplet. However, we should point out that a rough analysis of the relative intersystem crossing yield for 6F-6F, 6F-6H and 6F-ODA does not parallel the difference in decomposition rate exhibited in Fig. 2 for 6F-6F versus 6F-ODA and 6F-6H. While quantitative results are not available, our results seem to indicate that the substituent effect on the N-aryl ring $-(\text{CF}_3)_2\text{CPh}$ versus $-(\text{CH}_3)_2\text{CPh}$ or $-\text{OPh}$ not only alters the intersystem crossing yield but also has an effect on the subsequent reactions which proceed from the triplet state. Hence, any proposed photodegradation mechanism, such as depicted in scheme I (suggested⁵ as being consistent with our cumulative results to date for model and polymer systems having the N-arylimide moiety), must reflect the possibility that the R substituent group can affect both the properties of the excited state from whence products are formed, as well as subsequent reactions of any intermediates, such as a biradical comprised of an amidyl and a carbonyl radical, which may form. This will, of necessity, require positive identification of the spin multiplicity of the excited state leading to product formation. The rest of this paper will be directed toward identification of the effect of D-A properties on the photophysics and photochemistry of model N-arylphthalimides by clearly defining the spin multiplicity of the reactive states, the D-A effects on singlet emission and

singlet to triplet intersystem crossing yields of N-arylphthalimides, the energy levels of the singlet/triplet states of N-arylphthalimides, and the lifetimes of the reactive states.

B. Photophysics of Model N-Arylphthalimides

Absorption⁵ spectra of the N-arylphthalimides (see structures) shift to longer wavelengths as the D-A character increases. Furthermore, there is a measurable solvent effect on the fluorescence (first five entries in Table II) of each of these arylphthalimide models : The emission maxima shift to lower energies with increasing solvent polarity. We note that the N-aryl group seems to be essential for observation of the broad, substituent and solvent sensitive band in the fluorescence spectra of the N-arylphthalimides, since the fluorescence of N-alkylphthalimides (last two entries in Table II) is considerably blue shifted relative to the N-arylphthalimides and shows little solvent sensitivity. For example, the fluorescence maximum of PA-Pr in cyclohexane is at 410 nm while it is at 507 nm for PA-A. Furthermore, in contrast to PA-A, the fluorescence maximum of PA-Pr does not shift when the solvent is changed from cyclohexane to acetonitrile, although its intensity increases considerably. However, the fluorescence maximum of PA-A shifts from 507 to 552 nm in going from cyclohexane to acetonitrile, with a significant drop in intensity (Table II).

In order to provide additional insight into the photophysical nature of

N-arylphthalimides, we measured the fluorescence quantum yields (Φ_F) of the models in several different solvents and at excitation wavelengths from 250 nm to 350 nm (Table III) : We did not observe an excitation wavelength effect on Φ_F for any of the models or polymers examined. Moreover, the fluorescence was weak in all cases with the highest value of Φ_F being 8.3×10^{-3} for the model compound PA-CNA in cyclohexane. Upon close examination of the results in Table III, it can be seen that Φ_F diminishes for a given compound in more polar solvents and as the D-A character of the material is enhanced by electron donating substituents. In contrast to the results for the N-arylphthalimides, the 'isolated' phthalimide chromophore, as exemplified by PA-Pr, exhibits a marked increase in Φ_F from 5.2×10^{-5} to 1.0×10^{-3} in going from cyclohexane to acetonitrile. Solvent effects on the fluorescence of N-alkylphthalimides⁹ have been attributed to an increase in π, π^* character of the singlet state in solvents of increasing polarity. This seems quite reasonable to us. Obviously the same mechanism cannot be invoked for the N-arylphthalimides, or for the 6F-6F, 6F-6H, and 6F-ODA polymers since an increase in solvent polarity leads to a pronounced red shift in the position of the wavelength maximum and an accompanying decrease in the fluorescence intensity. This is best exemplified for 6F-6F in Table I since we were able to obtain fluorescence spectra in three solvents for this polymer.

From the results presented in this report, it seems obvious that D-A

effects are important to the singlet state behavior of N-arylphthalimides and aromatic based polyimides. One possibility is that N-arylimide groups have a twisted intramolecular charge transfer (TICT) state in which the N-aryl group acts as an electron donor and the phthalimide as an acceptor. Indeed, Frank et al.¹²⁻¹⁴ have previously suggested that the red shifted fluorescence of N-aryl polyimides originate from an intramolecular charge transfer complexes. While there is no direct proof that N-arylphthalimides exhibit a TICT state, they do have a very broad, weak, structureless fluorescence that is dependent on the donor/acceptor character of the interactive components and to solvent polarity, with a large Stokes shift from the onset of absorption, all characteristic of a TICT state.^{10,11} A compound forming a TICT state is also expected to have a very small singlet-triplet energy gap : Such is apparently the case for N-arylphthalimides. For example, the phosphorescence spectrum of PA-CNA at 77K (Fig. 3) in an ether-tetrahydrofuran glass, shows a 0-0 band at 402 nm, corresponding to a triplet energy, E_T , of 71.1 kcal. mol⁻¹. Such precise determination of the singlet energy, E_S , is not possible, since no vibrational structure is apparent in the fluorescence spectrum. However, taking the wavelength at which the fluorescence intensity is 10% of its value at λ_{max} as an approximation for the position of a 0-0 band, we obtain a value for E_S of ca. 72 kcal. mol⁻¹ in cyclohexane, and ca. 70 kcal. mol⁻¹ in acetonitrile, identical within experimental error to the triplet state energy obtained from the

phosphorescence experiments (Table IV). Similar results are obtained for the other model compounds and for 6F-6F and 6F-ODA (Table IV).

C. Spin Multiplicity of Reactive State :

Although we have presented data for N-arylphthalimides with respect to the effect of donor/acceptor substituents on the singlet state, we are ultimately interested in the triplet state, since preliminary quenching results suggest⁵ that photooxidative degradation of the phthalimide model compounds proceeds from the triplet state of the imide. Therefore, we have used laser flash photolysis to characterize the lowest triplet states of the polymers and model compounds in fluid solution. Laser flash photolysis (248 nm) of the models PA-CNA, PA-A, PA-CIA, PA-POA, and PA-MOA led to observation of transients with lifetimes of several microseconds in nitrogen saturated solvents. The transient lifetimes are greatly reduced in each case when the medium is air saturated. For example, the transient from PA-A in cyclohexane has a sharp maximum at ca. 300 nm and a broad maximum at ca. 520 nm (Fig. 4). The lifetime of this transient is ca. 12 μ s in a nitrogen purged solution and ca. 150 ns in the presence of air. The decay of the transient absorption is independent, within experimental error ($\pm 10\%$), of the monitoring wavelength indicating that it is due to a single species. The transient is readily quenched by cyclohexadiene ($E_T = 52.4$ kcal. mol⁻¹) with a quenching rate constant of 6.0×10^9 L.mol⁻¹ s⁻¹, and by

naphthalene ($E_T = 61 \text{ kcal. mol}^{-1}$) with a quenching rate constant of $2.9 \times 10^9 \text{ L.mol}^{-1} \text{ s}^{-1}$. Quenching with biphenyl ($E_T = 66 \text{ kcal. mol}^{-1}$) is much less efficient ($< 2 \times 10^7 \text{ L.mol}^{-1} \text{ s}^{-1}$). Formation of the transient can be sensitized by 351 nm excitation in the presence of benzophenone ($E_T = 69 \text{ kcal. mol}^{-1}$). Similar results are obtained for the other N-arylphthalimide model compounds, as well as the 6F-6F and 6F-ODA polymers (Fig. 5) in solution, although the lifetimes of transients from the polymers in dilute solution are much shorter (nitrogen saturated solution) than for the models (Table V). The short lifetimes for 6F-6F and 6F-ODA suggest the possibility of intra-chain self quenching in the polymers that cannot occur in dilute solutions of the models. The consequences of this self quenching phenomenon are the subject of another paper.

From the data obtained for all of the models and polymers investigated, there seems little doubt that the transient species arising from the excitation of N-arylphthalimides are due to triplet states of the phthalimides with energies between 61 and 66 kcal. mol⁻¹ (estimated from quenching studies at room temperature) or 67 - 71 kcal. mol⁻¹ (estimated from phosphorescence spectra at 77 K). In order to confirm this hypothesis, energy transfer experiments were carried out using β -carotene as the energy acceptor. β -Carotene is a convenient triplet indicator since it gives no transient absorption upon direct laser photolysis due to an inherently low intersystem crossing quantum yield.¹⁵ When PA-A is excited

in an excess of β -carotene, the PA-A transient is completely quenched and triplet β -carotene is produced¹⁵ with a characteristic absorption maxima around 520 nm (Fig. 6). The growth of the carotene triplet and the decay of PA-A transient absorption occur on essentially the same time scale indicating that the transient species produced on excitation of PA-A is a triplet : The 400 - 470 nm region in Figure 6 shows a bleaching effect due to the depletion of ground state β -carotene absorption.

In order to confirm the role of the triplet state in the photooxidative degradation of N-arylphthalimides, photolysis experiments^{4,5} (using a 450 W medium pressure mercury lamp) were conducted using triplet quenchers and sensitizers. The photolysis of PA-A ($\lambda_{ex} = 313$ nm, band pass filter) in the presence of naphthalene (triplet quencher) in methylene chloride, resulted in the quenching of the quantum yield for phthalic anhydride formation (Φ_{PA}) with a bimolecular quenching rate constant of 9.0×10^9 L. mol⁻¹ s⁻¹, well within the experimental error of the quenching constant obtained by laser flash photolysis. Competitive absorption by naphthalene at 313 nm was negligible for the concentrations employed. A comparable quenching rate constant was also obtained when cyclohexadiene was used. Similarly, it was found that the extent of photolysis, as monitored by GPC, of the 6F-6F polymer when irradiated with 313 nm radiation was considerably lower when carried out in the presence of the triplet quencher, 1,3-cyclohexadiene. After 7 hrs photolysis the reduction in the peak

maximum molecular weight (compared to polystyrene) was only 5 % compared to a reduction of 57 % in the absence of the quencher. When benzophenone was used as the triplet sensitizer for PA-A, Φ_{PA} was increased to a value of 5×10^{-3} , over an order of magnitude larger than obtained by direct photolysis in dichloromethane. This is consistent with the sensitization expected by benzophenone with a triplet energy of 69 kcal mol⁻¹, which is greater than either the estimated value of 67.6 kcal mol⁻¹ for PA-A in an ether - tetrahydrofuran glass at 77 K (Table IV) or that estimated from the triplet quenching results at room temperature. The quenching and sensitization results provide compelling evidence that the photooxidative degradation of the model compounds proceeds from the triplet state of N-arylphthalimides.

In order to provide a comparison of the intersystem crossing efficiencies of the model N-arylphthalimides, intersystem crossing quantum yields (Φ_{ISC}) were determined by laser flash photolysis relative to benzophenone ($\Phi_{ISC} = 1.0$). The Φ_{ISC} of the N-arylphthalimides decreases with increasing D-A character and with increasing solvent polarity as can be seen in Table VI (some of these values were extracted from previous papers⁶). These results also parallel the observation we obtained by emission and irradiation studies (most of which were reported in reference 5) reflecting the relative sensitivity of these compounds to photooxidative degradation: i.e., PA-CNA > PA-A > PA-POA > PA-MOA. However, as we

noted in the previous paper in this series⁵, substituent effects are probably prominent not only in increasing the quantum yield for intersystem crossing to the reactive triplet state, but also the subsequent reactions which occur. Interestingly, the triplet lifetime itself and hence the actual initial reaction from the triplet state is not altered appreciably by the substituent at the para position of the N-aryl ring, i.e., for example see entries for PA-A in cyclohexane in Table V. Since we find that the quantum yield for the products formed increases to a greater extent than the intersystem crossing yield upon substitution of an electron withdrawing group on the donor portion of the molecule, it seems probable that such substituents play a role in determination of the reactivity of intermediates other than by just enhancing intersystem crossing to the triplet state. There is little question that the correlation of photochemical reactivity of N-arylphthalimide moieties with donor substituents on the N-aryl group is important.

D. Conclusion

In this work, we have established the triplet multiplicity of the state responsible for the photochemistry of N-arylphthalimides and polyimides based on partially fluorinated polymers. The origin of the photochemical dependence of the D - A properties of the major components comprising the N-arylphthalimide repeat unit has been established quantitatively. A viable mechanism for the photolysis in an oxygen environment of partially

fluorinated polyarylimides has been postulated. Finally, the fluorescence emission of N-arylphthalimides from a TICT state seems feasible in view of the fluorescence solvent and substituent dependence, as well as the close proximity of the singlet - triplet energy levels in N-arylphthalimides.

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Table I. Fluorescence Maxima (nm) of the Polymers in Fluid Solution at Room Temperature

Polymer	Methylene Chloride	Tetrahydrofuran	Acetonitrile
6F-ODA	595	600	Not Soluble
6F-6H	575	580	very weak
6F-6F	525	530	542

Table II. Fluorescence Maxima (nm) of Phthalimide Model Compounds in Fluid Solution at Room Temperature.

Compound	Cyclohexane	Methylene Chloride	Acetonitrile
PA-CNA	471	493	512
PA-A	507	537	552
PA-CIA	509	540	555
PA-POA	555	606	610
PA-MOA	565	622	a
PA-Pr	410	410	410
PA-Me	406	410	410

^aWeak emission between 610 and 650 nm.

Table III. Quantum Yields^a of Fluorescence of Phthalimide Model Compounds in Fluid Solution at Room Temperature.

Compound	Cyclohexane	Methylene Chloride	Acetonitrile
PA-CNA	8.3×10^{-3}	4.6×10^{-3}	1.8×10^{-3}
PA-A	2.5×10^{-3}	5.6×10^{-4}	2.5×10^{-4}
PA-CIA	2.3×10^{-3}	4.9×10^{-4}	2.5×10^{-4}
PA-POA	3.5×10^{-4}	4.9×10^{-5}	2.8×10^{-5}
PA-MOA	3.5×10^{-4}	3.9×10^{-5}	1.8×10^{-5}
PA-Pr	5.2×10^{-5}	1.1×10^{-3}	1.0×10^{-3}
PA-Me	4.3×10^{-5}	9.4×10^{-4}	1.4×10^{-3}

^a relative to anthracene ($\Phi_F = 0.3$); $\pm 10\%$.

Table IV. Excited State Energies from Emission Spectra of N-Arylphthalimides and Polymers.

Compound	E_T^a , kcal. mol ⁻¹	E_S , kcal. mol ⁻¹	
		Cyclohexane	Acetonitrile
PA-CNA	71.1 ^b	(72.2) ^c	(69.6) ^c
PA-A	(67.6) ^c	(69.6) ^c	(65.4) ^c
PA-POA	(67.1) ^c	(64.0) ^c	(58.7) ^c
		Film	Tetrahydrofuran
6F-6F	71.5 ^b	(63.5) ^c	(72.2) ^c
6F-ODA	71.1 ^b	(66.9) ^c	d

^aEther-tetrahydrofuran glass at 77K.

^bEstimated from 0-0 band.

^cEstimated from onset of emission (wavelength at which intensity of corrected spectrum is 10% of maximum intensity).

^dNot determined

Table V. Properties of Transient Species Measured by Laser Flash Photolysis.

Compound	Solvent	λ_{max} nm	$\tau(N_2)$	$\tau(air)$
PA-CNA	acetonitrile	310,340,550	15.0 μ s	439ns
	methylene chloride	310,340,540	2.6 μ s	590ns
	cyclohexane	300,340,560	8.0 μ s	225ns
PA-A	acetonitrile	300,500	11.4 μ s	116ns
	methylene chloride	310,340,500	10.8 μ s	430ns
	cyclohexane	300,330,520	12.4 μ s	155ns
PA-POA	acetonitrile	310,350,540	8.6 μ s	88ns
	methylene chloride	310,350,560	9.0 μ s	258ns
	cyclohexane	300,345,540	9.3 μ s	91ns
6F-ODA	methylene chloride	310,350,520	730ns	135ns
6F-6F	methylene chloride	310,340,520	650ns	127ns

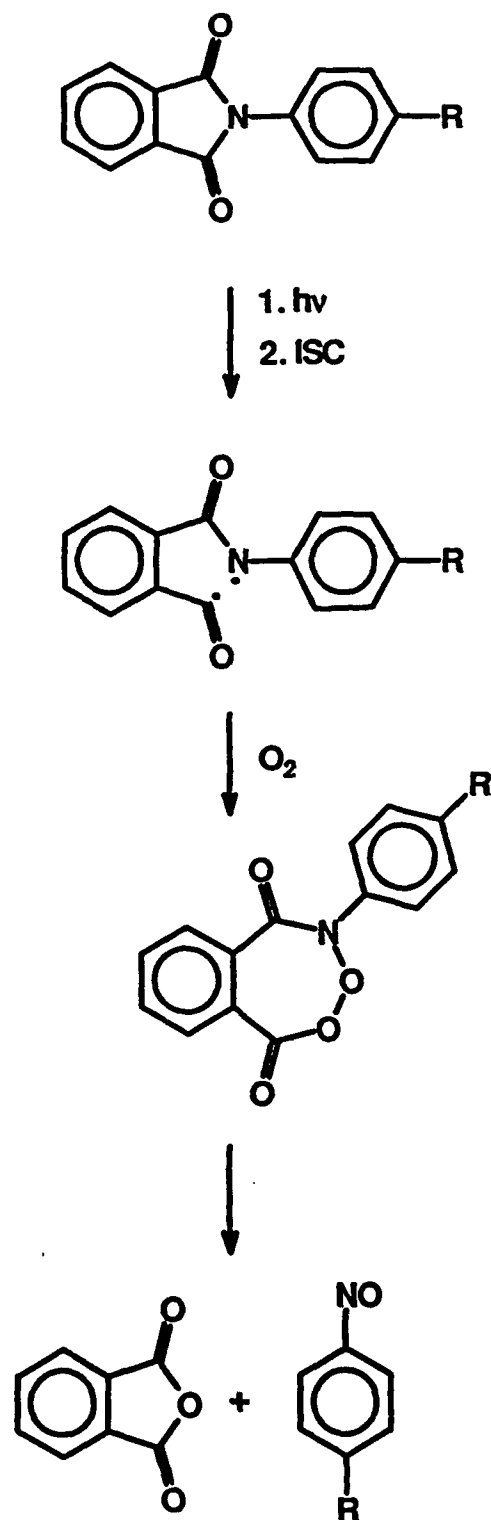
Table VI. Quantum Yields^a of Intersystem Crossing in Fluid Solution of N-Arylphthalimides.

Compound	Cyclohexane	Methylene chloride	Acetonitrile
PA-CNA	0.64	0.16	0.06
PA-A	0.17	0.04	0.03
PA-POA	0.06	0.02	0.02
PA-MOA	0.08	0.02	0.02

^a±15%; Determined relative to benzophenone ($\Phi_{ISC} = 1.0$)

FIGURE CAPTIONS

- Figure 1. Absorption spectra of free standing polyimide films cured at 300°C. (a) 6F-6F; (b) 6F-6H; (c) 6F-ODA.
- Figure 2. Plot of molecular weight of 6F-ODA (O), 6F-6F (□) and 6F-6H (★) films at peak maxima (compared to polystyrene) versus photolysis time with unfiltered medium pressure Hg lamp in air.
- Figure 3. Phosphorescence spectrum ($\lambda_{\text{ex}} = 250 \text{ nm}$) of PA-CNA in ether-tetrahydrofuran glass at 77K.
- Figure 4. Transient absorption spectra ($\lambda_{\text{ex}} = 248 \text{ nm}$) of PA-A in nitrogen-saturated cyclohexane solution at different time intervals after the flash.
- Figure 5. Transient absorption spectra ($\lambda_{\text{ex}} = 248 \text{ nm}$) of 6F-ODA in nitrogen-saturated methylene chloride solution at different time intervals after the flash.
- Figure 6. Absorption spectra ($\lambda_{\text{ex}} \approx 248 \text{ nm}$) at different time intervals after the flash, of β -carotene triplet generated by energy transfer from the PA-A triplet in cyclohexane.



Scheme I Mechanism for the Photolysis of N-arylphthalimides

Fig. 1

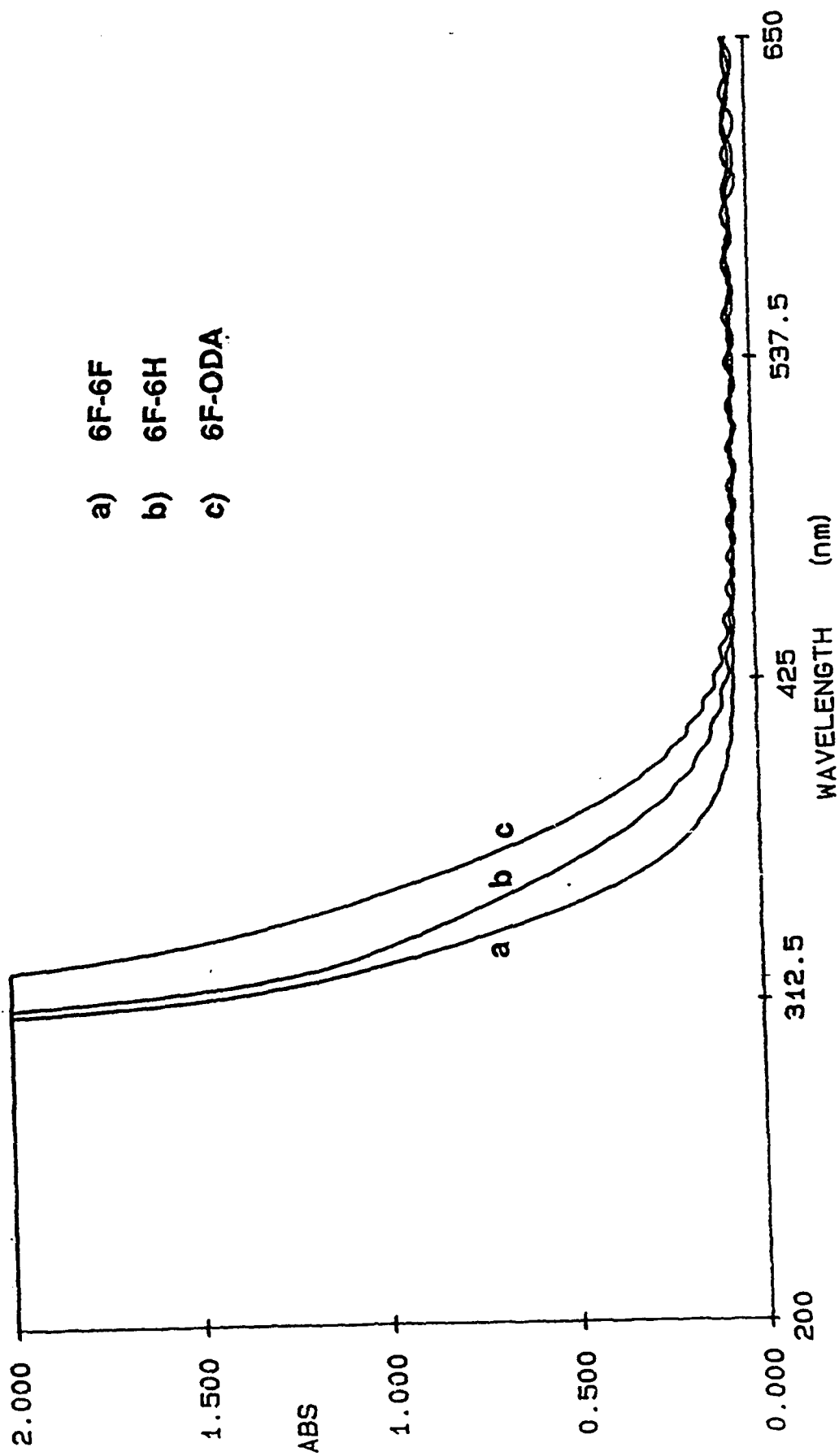


Fig. 2

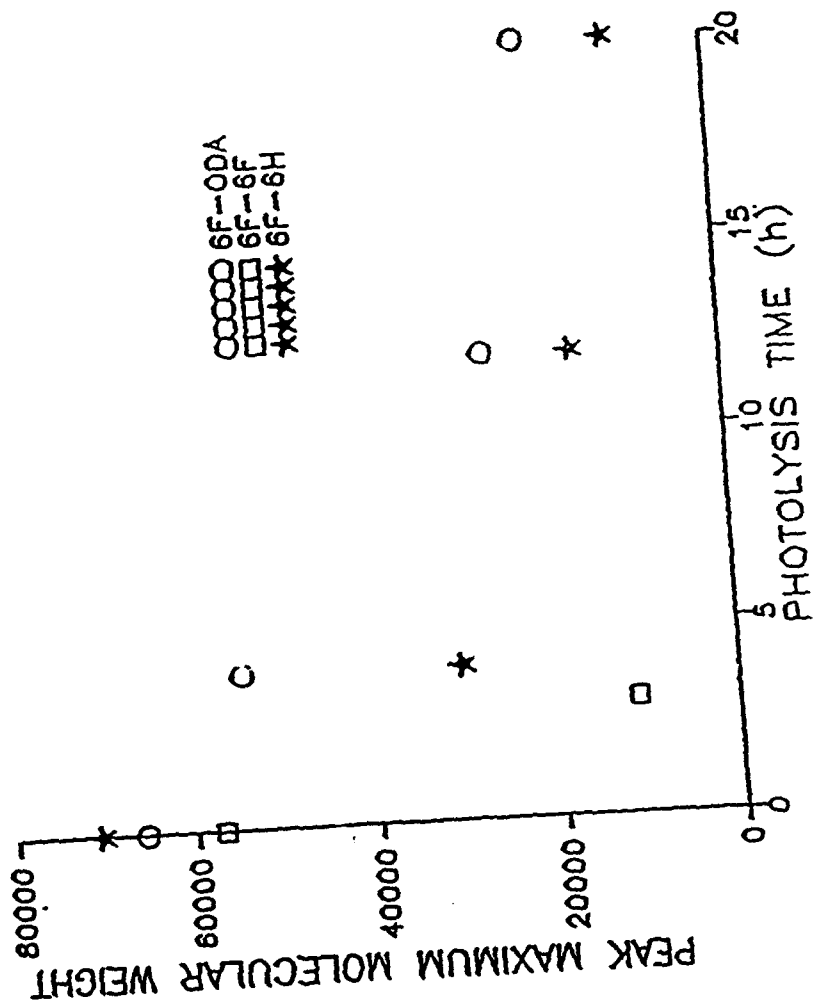


Fig. 3

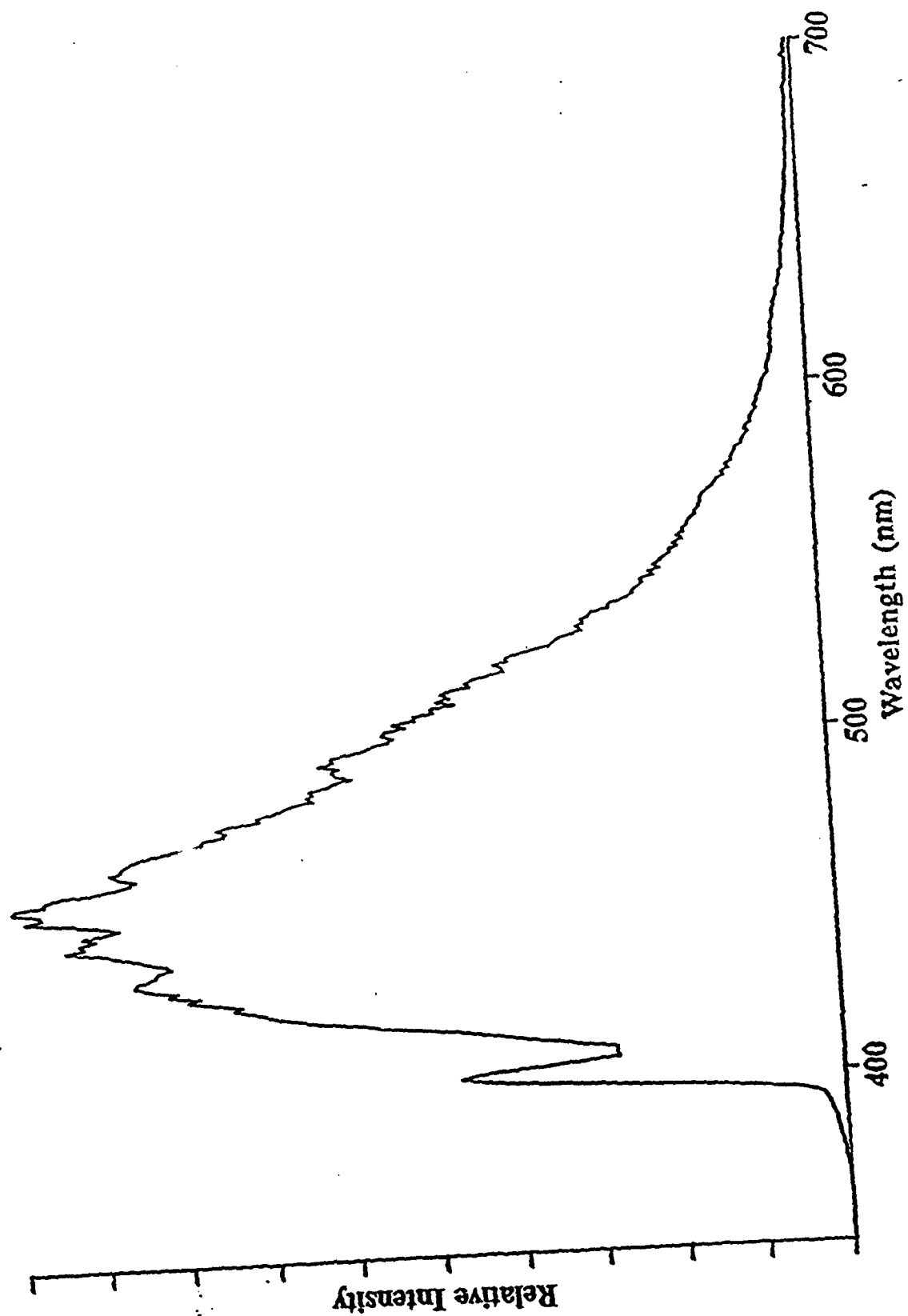


Fig. 4

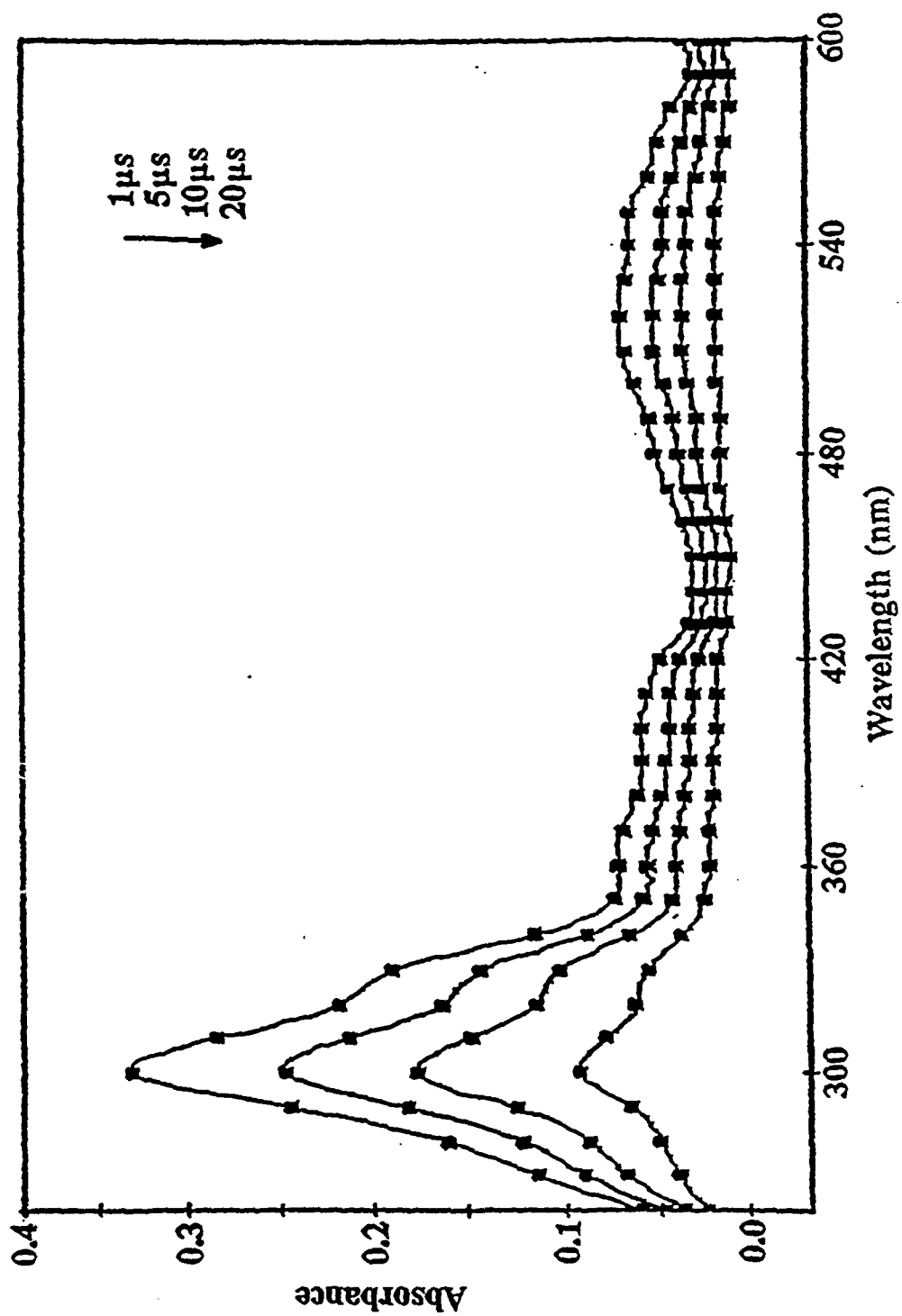


Fig. 5

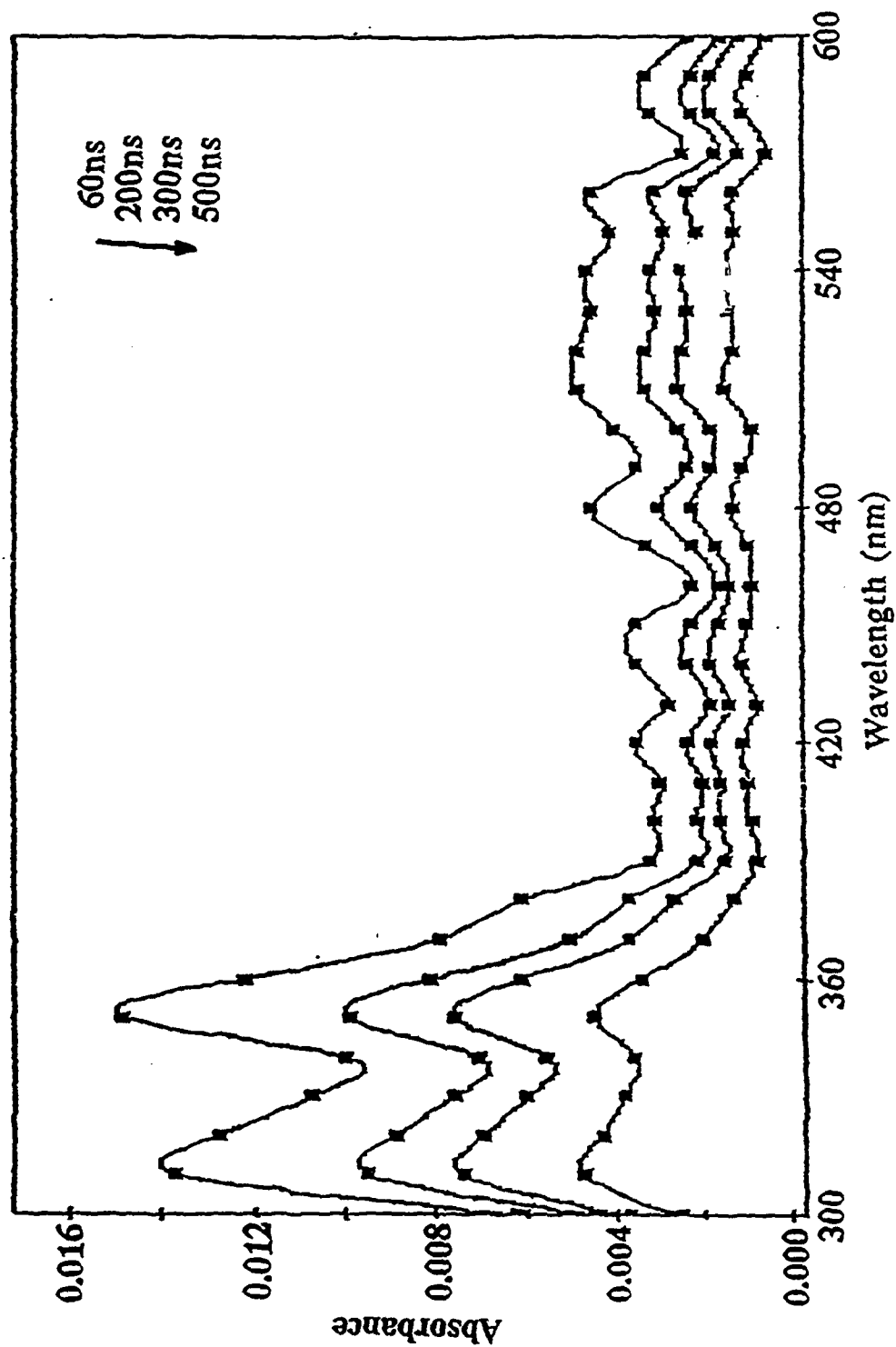


Fig. 6

